



a molecule has a mean velocity \bar{v}
 During "t", the particle travels across $V = \pi d^2 \bar{v} t$.
 Mean free path $\lambda = \frac{\bar{v} t}{\pi d^2 \bar{v} t \times n_V}$ particle concentration.

Consider now the displacement of the other molecules.
 relative velocity: $\bar{v}_r^2 = |\bar{v}_1 - \bar{v}_2|^2 = \bar{v}_1^2 + \bar{v}_2^2 - 2\bar{v}_1\bar{v}_2$
 uncorrelated velocities $\Rightarrow \bar{v}_1\bar{v}_2 = 0$
 $\Rightarrow \bar{v}_r = \sqrt{2} \bar{v}$ (statistically $\bar{v}_1 = \bar{v}_2 = \bar{v}$)

① Ideal gas $\Rightarrow PV = nRT$
 and volume density of particles: $n_V = \frac{n \times N_A}{V}$
 $\Rightarrow n_V = \frac{N \times P}{RT}$
 \Rightarrow Mean free path: $\lambda = \frac{\bar{v} t}{\bar{v}_r t \pi d^2 n_V}$
 $\Rightarrow \lambda = \frac{RT}{\sqrt{2} \pi d^2 N_A P}$ ↳ Avogadro number

Porosity of some arrangements

① elementary volume

$$\epsilon = \frac{S_s}{S_s + S_d} = \frac{\pi d^2 / 4}{(d+a)^2} = \frac{\pi}{4(1 + \frac{a}{d})^2}$$

$$\epsilon_{max} \text{ when } a = 0 \Rightarrow \epsilon_{max} = \frac{\pi}{4} = 0.79.$$

②
$$\epsilon = \frac{S_s}{S_s + S_d}$$

$$\Rightarrow S_s + S_d = \frac{3\sqrt{3}}{2} (d+a)^2 \text{ and } S_d = 3 \times \frac{\pi d^2}{4}$$

$$\Rightarrow \epsilon = \frac{\pi}{2\sqrt{3} (1 + \frac{a}{d})^2} \Rightarrow \epsilon_{max} (a=0) = \frac{\pi}{2\sqrt{3}} \approx 0.91$$

③ Elementary volume

$$V_s + V_d = d^3$$

$$V_s = V_s + V_d - V_d = d^3 - 8 \times (\frac{4}{3} \pi (\frac{d}{2})^3) \times \frac{1}{8}$$

$$= d^3 [1 - \frac{\pi}{6}]$$

$$\Rightarrow \epsilon = \frac{V_s}{V_s + V_d} = 1 - \frac{\pi}{6} \approx 0.48$$

Specific surface.

$$S_s = \frac{S_s}{V_s + V_d} \quad \left| \begin{array}{l} V_s + V_d = d^3 \\ S_d = 8 \times \frac{1}{8} \times 4\pi (\frac{d}{2})^2 = \pi d^2 \end{array} \right.$$

$$\Rightarrow S_s = \frac{\pi}{d}$$